

DRAFT UGANDA STANDARD

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Commercial blasting explosives — Specification — Part 3: Part Ammonium Nitrate for explosives



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DUS 1756-3: 2017

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Foreword

Uganda National Bureau of Standards (UNBS) is a parastatal under the Ministry of Trade, Industry and Cooperatives established under Cap 327, of the Laws of Uganda, as amended. UNBS is mandated to coordinate the elaboration of standards and is

- (a) a member of International Organisation for Standardisation (ISO) and
- (b) a contact point for the WHO/FAO Codex Alimentarius Commission on Food Standards, and
- (c) the National Enquiry Point on TBT Agreement of the World Trade Organisation (WTO).

The work of preparing Uganda Standards is carried out through Technical Committees. A Technical Committee is established to deliberate on standards in a given field or area and consists of key stakeholders including government, academia, consumer groups, private sector and other interested parties.

Draft Uganda Standards adopted by the Technical Committee are widely circulated to stakeholders and the general public for comments. The committee reviews the comments before recommending the draft standards for approval and declaration as Uganda Standards by the National Standards Council.

The committee responsible for this document is Technical Committee UNBS/TC 5, *Chemicals and Environment*

US 1285 consists of the following parts, under the general title *Commercial blasting explosives — Specification*

- — *Part 1: Emulsion Explosive*
- — *Part 2: Ammonium Nitrate – Fuel Oil Explosive*
- — *Part 3: Ammonium Nitrate - For explosives*
- — *Part 4: Sodium Nitrate - For explosives*

Commercial blasting explosives — Specification — Part 3: Part Ammonium Nitrate for explosives

1 Scope

This Draft Uganda Standard specifies requirements, sampling and test methods for Ammonium Nitrate intended primarily for use in explosives.

NOTE The requirements of this standard should be read in conjunction with the Act of 1964 cap 309 and/or other applicable regulations.

2 Normative references

The following referenced documents referred to in the text in such a way that some or all of their content constitutes requirements of this document. For dated references, only the edition cited applies. For undated references, the latest edition of the referenced document (including any amendments) applies.

DUS 1776, *light metal in hazardous location at mines — Guidelines for use*

DUS 1757, *Commercial blasting explosives — Terminology*

DUS ISO 7010, *Graphical symbols — Safety colours and safety signs — Registered safety signs*

3 Terms and definitions

For the purposes of this document, the terms and definitions given in DUS 1757 and the following apply

3.1

anti-caking agents

an additive placed in powdered or granulated materials, to prevent the formation of lumps (caking) and for easing packaging, transport, and consumption.

4 Grades

The material shall be of any of the following three grades namely;

- a) Grade 1 — for Ammonium Nitrate Fuel Oil (ANFO);
- b) Grade 2 — for slurry type explosives; or
- c) Grade 3 — for Nitroglycerine based explosives.

5 Requirements

5.1 Physical requirements

5.1.1 The material shall be in the form of prills, lumps or crystals, as agreed to between the purchaser and the supplier, and colourless or showing only a brownish tint. It shall be free from grit, foreign matter and visible impurities.

5.1.2 The different grades of material shall conform to the sizes given in Table 1.

Table 1 — Grade sizes

Grade	Size
	mm
Grade 1	0.65 – 2.8
Grade 2	2.9 – 100
Grade 3	As agreed to between the purchaser and the supplier.

5.2 Anti-caking agents

Where anti-caking agents are added they shall be compatible with the explosives and shall not cause undue frothing in the manufacturing process of explosives and the quantity and nature of the agent shall be declared by the supplier

5.3 Chemical requirements

All permitted ammonium nitrate intended primarily for use in explosives, in addition to meeting the general requirements stipulated in clause 5.1 above, and shall comply with chemical specification as specified in the table 2 below.

Table 2 — Chemical requirements of Ammonium Nitrate

NO.	Characteristic/ parameter	Grade			Test method
		%			
		Grade 1	Grade 2	Grade 3	
1	Moisture, percent by mass, (Max)	0.3	3.0	0.1	Annex A
2	Matter insoluble in water, percent by mass, (Max)	0.3	0.05	0.01	Annex B
3	Non-volatile matter, percent by mass, (Max)	0.2	-	0.1	Annex C
4	Chlorides (as Cl), percent by mass, (Max)	-	0.05	0.006	Annex D
5	Nitrites (as Ammonium Nitrite), percent by mass, (Max)	-	0.02	0.0005	Annex E
6	Acidity (as Nitric acid), percent by mass, (Max)	-	0.01	0.01	Annex F
7	Iron (as Fe), percent by mass, (Max)	-	0.002	0.002	Annex G

8	Calcium (as calcium nitrite), percent by mass, (Max)	-	0.5	-	Annex H
9	pH of 10 percent aqueous solution	≥ 4	5 - 6	4 -7	Annex I
10	Organic matter, percent by mass, (Max)	0.5	Nil	0.1	Annex J
11	Grit, percent by mass, (Max)	-	Nil	Nil	Annex K
12	Purity (on dry basis), percent by mass, (Min)	99.2	99.0	99.8	Annex L
13	Metals and compounds of metals other than alkali metals, calcium and iron, percent by mass, (Max)	-	-	0.02	Annex M
14	Pyridine, percent by mass, (Max)	-	-	Nil	Annex N
15	Thiocyanate (as ammonium thiocyanate), percent by mass, (Max)	-	-	Nil	Annex O
16	Prussian blue and allied compounds (as Prussian blue), percent by mass, (Max)	-	-	Nil	Annex P
17	Oil absorption, percent by mass, (Min)	7	-	-	Annex Q
18	Total nitrogen, percent by mass, (Min)	34.5	-	-	Annex R
19	Sulphates (as ammonium sulphate), percent by mass, (Max)	-	-	0.03	Annex T

5.4 Bulk density

When determined in accordance with Annex R, the bulk density of the Ammonium Nitrate for the respective grades shall be as that given in table 3 below

Table 3— Bulk density of ammonium nitrate

Grade	Bulk density g/cm ³
Grade 1	0.65 – 0.85
Grade 2	-
Grade 3	0.92 – 1.02

6 Inspection and sampling

6.1 General

Owing to the dangerous nature of ammonium nitrate intended primarily for use in explosives, extreme care and caution shall be exercised during the inspection, sampling and testing of ammonium nitrate. All safety precautions and procedures laid down in the relevant national legislations shall be strictly followed.

6.2 Inspection

Inspect the ammonium nitrate for compliance with the requirements of clause 5.

6.3 Sampling

In a single consignment of the material all the containers of the same, type and size and drawn from the same batch of manufacture shall constitute a lot. If a consignment is known to consist of different batches of manufacture or of different types and sizes of containers, the containers belonging to the same batch type and size shall be grouped together and each such group shall constitute a separate lot.

The number of containers to be selected at random from lots of different sizes shall be in accordance with Table 4.

Table 4 — Sample sizes to be considered

Lot size	Sample size
Up to 15	3
16 to 25	4
26 to 50	5
51 to 100	7
101 and above	10

7 Packaging and labelling

7.1 Packaging

The material shall be packed in polyethylene lined bags or in such other containers as agreed to between the purchaser and the supplier. The polyethylene sheet shall not be less than 0.05 mm thick

7.2 Labelling

7.2.1 Each package shall also be marked with appropriate safety symbol as specified in US ISO 7010.

7.2.2 Each package shall be legibly and indelibly marked with the following information:

- a) name and grade of the material;
- b) manufacturers' name or his recognized trade-mark, if any;
- c) year of manufacture and expiry date;
- d) lot/ batch number;
- e) country of origin; and
- f) gross and net weight.

In addition to the above, the following cautionary note shall also appear on the label in red ink: "HIGHLY EXPLOSIVE. HANDLE WITH UTMOST CARE. DO NOT BRING ANY SPARK OR FLAME NEAR THE PACKAGE. DO NOT STORE, TRANSPORT OR STOCK WITH THE DETONATORS."

Annex A (normative)

Determination of moisture using air-oven method

A.1 Procedure

- A.1.1** Weigh rapidly 10 g of the material and transfer it to a weighing dish with a close fitting lid.
- A.1.2** Close the dish and weigh it accurately.
- A.1.3** Lift the lid and expose the material in an oven at $100\text{ }^{\circ}\text{C} \pm 2\text{ }^{\circ}\text{C}$ for 2 h.
- A.1.4** Then close the lid, remove the dish to a desiccator for about half an hour to cool and reweigh it.
- A.1.5** Repeat the operation till a constant mass is obtained.

A.2 Calculation

$$\text{Moisture, percent by mass} = 100 \times \frac{M_1 - M_2}{M_1 - M}$$

1

where

M_1 is the mass in g of the material and the dish before heating;

M_2 is the mass in g of the material and the dish after heating;

M is the mass in g of empty dish.

Annex B (normative)

Determination of matter insoluble in water

B.1 Procedure

- B.1.1** Take 100 g of the sample in a 500-ml beaker and dissolve in about 200 ml of hot water.
- B.1.2** Filter the residue through What-man No. 40 filter paper.
- B.1.3** Wash the residue with water till free from all soluble matter.
- B.1.4** Dry in an oven maintained at $100\text{ }^{\circ}\text{C} \pm 2\text{ }^{\circ}\text{C}$.
- B.1.5** Brush out the residue in a weighed silica crucible and weigh.
- B.1.6** Deduct the quantity of the declared anti-caking agent insoluble in water, if any, from the dried residue.

B.2 Calculation

$$\text{Matter insoluble in water, percent by mass} = 100 \times \frac{M_1}{M} \quad 2$$

where

M_1 is the mass in g of the dried residue, and

M is the mass in g of the material taken for the test.

Annex C (normative)

Determination of non-volatile matter

C.1 Procedure

C.1.1 Heat a porcelain or silica basin with cover about 6 cm in diameter in an oven at $100\text{ }^{\circ}\text{C} \pm 2\text{ }^{\circ}\text{C}$ for half an hour.

C.1.2 Cool it in a desiccator and weigh.

C.1.3 Place about 10 g of the material in the dish, moisten the ammonium nitrate with sulphuric acid, replace the cover and weigh accurately.

C.1.4 Uncover the dish and heat the sample initially over low flame till the fume subsides and finally over strong flame having a temperature of $800\text{ }^{\circ}\text{C} \pm 20\text{ }^{\circ}\text{C}$ for 1 h.

C.1.5 Remove the burner and replace the cover, cool the basin and the cover in a desiccator to room temperature and weigh.

C.1.6 Repeat the operation till constant mass is obtained.

C.2 Calculation

Calculate the non-volatile matter as follows:

Non-volatile matter, percent by mass $= \frac{100 \times (M_3 - M_1)}{M_2 - M_1}$

where

M_1 mass in g of the empty dish and the cover;

M_2 mass in g of the dish and the cover with the sample taken; and

M_3 mass in g of the dish and the cover with the sample after heating.

Annex D (normative)

Determination of chlorides

D.1 Outline of the Method

Chloride is completely precipitated as silver chloride by the addition of an excess of standard silver nitrate solution and the excess is back titrated with ammonium thiocyanate.

D.2 Reagents

- 1) Silver Nitrate Solution – 0.1 N;
- 2) Ammonium Thiocyanate — 0.1 N;
- 3) Concentrated Nitric Acid;
- 4) Ferric Alum Indicator — 40 %; and
- 5) Nitrobenzene

D.3 Procedure

D.3.1 Weigh about 10 g of the sample accurately in a 250-ml conical flask. Dissolve with heating in 100 ml of water.

D.3.2 Add 5 ml of concentrated nitric acid followed by 25 ml silver nitrate solution through a pipette.

D.3.3 Shake for two minutes and titrate against 0.1 N thiocyanate using 2 ml of ferric alum indicator and 5 ml of nitrobenzene till there is a persistent colour change.

D.3.4 Run a blank with all the reagents and water but without the sample.

D.4 Calculation

$$\text{Chlorides, percent by mass} = (A - B) \times N \times \frac{35.45}{1000} \times \frac{100}{M}$$

4

where

- A* volume of ammonium thiocyanate consumed for blank titration;
- B* volume of ammonium thiocyanate consumed for sample titration;
- N* normality of ammonium thiocyanate solution; and
- M* mass in g of sample taken for the test.

Annex E (normative)

Determination of nitrite

E.1 General

Two methods are prescribed, namely, colorimetric method and volumetric method. Colorimetric method is recommended for routine analysis. In case of dispute, volumetric method shall be the referee method.

E.2 Colorimetric method

When dilute solutions of sulphanilic acid and α -naphthylamine are acted upon by nitrous acid, a red colouration is produced which may be used for the colorimetric determination of nitrites. The full colour does not appear at once. A satisfactory colour comparison can be made after 10 min - 15 min provided the sample and standard are treated in the same manner.

E.2.1 Reagents

E.2.1.1 Sulphanilic acid reagent (solution 1) — Dissolve 0.6 g of sulphanilic acid in 100 ml of 20 % (v/v) dilute hydrochloric acid.

E.2.1.2 α -Naphthylamine reagent (solution 2) — Dissolve 0.48 g of α -naphthylamine in 100 ml of 20 percent (v/v) dilute hydrochloric acid.

E.2.1.3 Standard nitrite solution — Dissolve 4.93 g of sodium nitrite in water and dilute with deionized water. Dilute 10 ml of standard solution to one litre with deionized water.

E.2.1.4 Sodium acetate (2 M) — Dissolve 16.4 g of anhydrous sodium acetate in water and dilute to 100 ml with water.

E.2.2 Procedure

E.2.2.1 Place the quantity of the sample containing about 0.03 mg or less of nitrite in a 50 ml volumetric flask.

E.2.2.2 Add 1 ml of sulphanilic acid reagent, mix well and allow to stand at least for 3 minutes but not more than 10 minutes at room temperature in diffused light. Introduce 1 ml of α -naphthylamine reagent and 1 ml of 2 M sodium acetate solution to act as buffer (pH 2 to 2.5).

E.2.2.3 Dilute to volume and mix well. Use a photoelectric colorimeter with a green filter and measure transmittance. Construct a reference curve with known nitrite standards, preferably in the nitrite-nitrogen range of the sample and measure the concentration of the nitrite in the sample solution from the reference curve.

E.2.3 Calculation

The concentration of nitrite is directly read from the reference curve and calculated as ammonium nitrite.

E.3 Volumetric method

E.3.1 Reagents

- 1) Potassium permanganate solution — N/50.
- 2) Sulphuric acid

E.3.2 Procedure

E.3.2.1 Weigh accurately about 30 g of sample and transfer to a 100-ml volumetric flask and dissolve with shaking in 40 ml of water.

E.3.2.2 Make up the volume to the mark with water. Stopper the flask and shake well. Take up the solution in a burette.

E.3.2.3 Take 2.5 ml of freshly prepared N/50 potassium permanganate solution in a conical flask. Add 5 ml of sulphuric acid (1:1)

E.3.2.4 Dilute to 40 ml with water and heat to around 60°C and titrate against sample solution taken in the burette till the pink colour disappears.

E.3.1 Calculation

1 ml of 1 N potassium permanganate = 0.0320 g of ammonium nitrite.

Annex F (normative)

Determination of acidity of nitric acid

F.1 Outline of the method

The dried material is dissolved in neutral acetone, water added to the solution and then titrated with sodium hydroxide solution.

F.2 Reagents

- 1) **Acetone** — Neutralized to methyl red indicator
- 2) **Standard Sodium Hydroxide Solution** — 0.01 N.
- 3) **Methyl Red Indicator** — Dissolve 1 g of methyl red (water soluble) in water and dilute the solution to 1 litre

F.3 Procedure

F.3.1 Add about 2 to 10 g of the dried material weighed accurately to 0.01 g, to 100 ml of acetone. Shake the mixture vigorously for several minutes until solution is complete.

F.3.2 Add 100 ml of water to solution with constant stirring. Boil off the acetone on a water bath, cool, filter through a filter paper (Whatman No. 1 or its equivalent), wash with two 10 ml portions of water and titrate the filtrate with standard sodium hydroxide solution using methyl red as indicator.

F.3.3 Run a blank determination using the same quantities of reagents

F.4 Calculation

$$\text{Acidity (as nitric acid), percent by mass} = \frac{6.3 (V_2 - V_1) N}{M}$$

5

where

V_1 = volume in ml of standard sodium hydroxide solution required by the blank,

V_2 = volume in ml of standard sodium hydroxide solution required by the material,

N = normality of standard sodium hydroxide solution, and

M = mass in g of the dried material taken for the test.

Annex G (normative)

Determination of iron

G.1 Outline of the method

Ferric (but not ferrous) iron reacts with thiocyanate to give a series of intensely red coloured complexes which remain in true solution. In the colorimetric determination, excess of thiocyanate should be used since this increases the intensity and also stability of the colour. Strong acid (hydrochloric or nitric acid concentration 0.052 to 0.05 M) should be present to suppress the hydrolysis of ferric iron.

G.2 Reagents

- 1) **Potassium Thiocyanate Solution** — Dissolve 20 g of potassium thiocyanate in 100 ml of water.
- 2) **Concentrated Nitric Acid**
- 3) **Standard Solution of Iron (Ferric)** — Dissolve 0.864 g of Ferric Ammonium Sulphate in water, add 10 ml of concentrated nitric acid and dilute it to one litre. 1 ml of this solution is equivalent to 0.1 mg of iron (as Fe).

G.3 Procedure

G.3.1 Dissolve accurately 5 g of the material in 1:1 water/nitric acid mixture and evaporate to nearly dryness to expel excess of acid. Dilute slightly with water, oxidize any ferrous ion to the ferric state with dilute potassium permanganate solution (2g/l), adding slowly dropwise until a slight pink colour remains after stirring well. Make up the solution to 250 ml or other suitable volume.

G.3.2 Place 50 ml of the solution in a Nessler cylinder, add 5 ml of thiocyanate solution and 2 to 4 ml of 4 N nitric acid. Add the same amounts of reagents to 50 ml of water contained in a similar Nessler cylinder and run in standard iron solution from burette (use long glass rod with flattened end for mixing) until colours are matched. Note the exact volume (X ml) of the standard iron solution added.

G.3.3 Repeat determination using 50 - X ml of water. Comparison of the standard and unknown should be made soon after preparation since the colour may fade on standing for a long time.

G.3.4 For estimating the concentration of ferric iron more accurately, the use of photoelectric colorimeter may be resorted to. A filter showing maximum transmission at or near 480 nm should be used. Standard solutions of ferric iron which are in the range of the concentration of ferric iron likely to be present in the sample are prepared and the transmittance of each such solution measured with photoelectric colorimeter.

G.3.5 A reference curve is then prepared by plotting the concentration versus transmittance. Now the transmittance of the sample solution is also measured and from the reference curve, the concentration of ferric iron as $[\text{Fe}(\text{NO}_3)_3]$ is directly read.

G.4 Calculation

$$\text{Ferric iron as } [\text{Fe}(\text{NO}_3)_3] \text{ percent by mass} = \frac{m}{M} \times 100$$

6

where

m = mass in g of ferric iron [as Fe (NO₃)₃] as found out from the reference curve, and

M = mass in g of the material taken for the test.

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Annex H (normative)

Determination of calcium

H.1 Outline of the method

Calcium, like some other metal ions, is complexed by Disodium Dihydrogen Ethylene Diamine Tetra-Acetate (EDTA) by choosing the proper pH and indicator. Using calcon as indicator, calcium can be estimated in the presence of magnesium since magnesium is precipitated quantitatively as magnesium hydroxide in the pH range of the reaction

H.2 Reagents

- 1) **Diethylamine**
- 2) **Calcon Indicator** — prepared by dissolving 0.2 g of the dyestuff in 50 ml of methyl alcohol.
- 3) **Standard Disodium Dihydrogen Ethylenediamine Tetra Acetate (EDTA) Solution** — Weigh 3.772 5 g of EDTA, dissolve in water free from polyvalent ions (distilled water for this purpose should be passed through a column of cation exchange resins in the sodium form) and make up to a volume of 1 000 ml. Concentration of the resulting solution will be 0.01 M. (If the EDTA is not of analytical reagent grade, then the solution prepared needs to be standardised against standard zinc chloride solution or magnesium chloride solution which is prepared from analytical reagent grade material or corresponding metal pellets of analytical reagent grade).
- 4) **Sample solution** — Weigh about 1 g of the material accurately, dissolve in water from free polyvalent ions, filter and make up to 100 ml.

H.3 Procedure

H.3.1 Pipette out 10 ml of the sample solution (see A.9.2.4) into 250 ml conical flask.

H.3.2 Add about 40 ml of water and 5 ml of diethylamine, giving the mixture a pH of about 12.5. Under these conditions, magnesium, if present, precipitates quantitatively as the hydroxide.

H.3.3 Add four drops of calcon indicator to the solution and titrate with standard EDTA solution with shaking until the colour changes from pink to a pure blue.

H.4 Calculation

$$\text{Calcium [as Ca (NO}_3\text{)}_2\text{], percent by mass} = \frac{1.64 * VM}{A}$$

7

where;

V = volume in ml of standard EDTA solution,

M = molarity of standard EDTA solution, and

A = mass in g of material taken for the test.

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**Annex I
(normative)**

Determination of pH

Dissolve 10.0 g of the material in 100 ml of water and measure the pH of the solution by means of a suitable pH meter using glass electrodes.

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Annex J (normative)

Determination of organic matter

J.1 Procedure

J.1.1 Weigh about 50 g of the sample accurately in a tall sintered-glass grade 2 thimble with a small hole at the top of wall for tying a small piece of wire.

J.1.2 Put the thimble gently in the soxhlet extraction chamber with the help of the wire.

J.1.3 Take 200 ml of the suitable solvent (dichloromethane, diethyl ether, acetone, N-butyl acetate, etc) in the flask containing a few glass beads. Fit the apparatus on a boiling water bath and continue the extraction for three hours.

J.1.4 Remove the condenser and take out the thimble.

J.1.5 Replace the condenser and continue boiling till majority of the solvent is recovered in the flask to a tared flat bottom glass basin with cover. Rinse the flask twice with the suitable solvent and transfer in the basin.

J.1.6 Put the basin in a boiling water bath at $95\text{ }^{\circ}\text{C} \pm 2\text{ }^{\circ}\text{C}$ for two hours.

J.1.7 At the end of this period, replace the cover, cool the basin and the cover in a desiccator to room temperature and weigh. Repeat the drying operation till constant mass is obtained.

NOTE When required, the manufacturers shall inform about the name of the solvent to be used.

J.2 Calculation

$$\text{Organic matter, percent by mass} = \frac{M_2}{M_1} * 100$$

8

where

M_2 = mass in g of the residue after complete evaporation of the solvent extract, and

M_1 = mass in g of the material taken for test.

Annex K (normative)

Determination of grit

K.1 Procedure

K.1.1 Weigh about 50 g of the material into a 500-ml beaker, dissolve in adequate quantity of water and allow to stand until the insoluble material settles.

K.1.2 Decant the supernatant liquor and filter the rest using a filter paper. Ignite the filter paper with the residue in a muffle furnace. Boil the residue for about 30 minutes with 20 ml aqua regia.

K.1.3 After cooling the mixture, drain off the aqua regia and wash the thoroughly with water.

K.1.4 Dry the residue in an air oven and sieve it through 125 µm IS-sieve.

K.1.5 Weigh the portion retained on the sieve and test for soda glass scratching. If it scratches soda glass then report it as percent grit.

K.2 Calculation

$$\text{Grit, percent by mass} = \frac{M * 100}{M_1}$$

9

Where;

M = mass in g of the residue retained on 125 µm IS-sieve, and

M₁ = mass in g of the material taken for test

Annex L (normative)

Determination of purity on dry basis

L.1 Outline of the method

Ammonia from ammonium nitrate (NH_4NO_3) reacts quantitatively with excess of formaldehyde (HCHO) to produce nitric acid (HNO_3).



This nitric acid can be titrated against standard alkali

L.2 Reagents

- 1) **Sodium Hydroxide** — 0.2 N.
- 2) **Hydrochloric Acid** — 0.2 N.
- 3) **Phenolphthaline** — 1 % in alcohol.
- 4) **Formaldehyde 1:1** — Mix equal volumes of AR grade formaldehyde (37 percent m/m) and water. Add few drops of phenolphthaline indicator and titrate against standard alkali to its lightpink end point

L.3 Procedure

- L.3.1 Weigh exactly 0.4 g of sample into a 250 ml conical flask and dissolve in 20 ml of water.
- L.3.2 Add 50 ml of formaldehyde 1:1 reagent and warm the flask.
- L.3.3 Cool to room temperature and titrate against standard alkali till permanent light pink colour appears.
- L.3.4 Drain another 2 ml of alkali and note down the volume (V_1).
- L.3.5 Keep it for 10 minutes and titrate against standard hydrochloric acid using a microburette till the solution becomes colourless and note down the volume (V_2).
- L.3.6 Repeat the same experiment till concordant values are obtained.

L.4 Calculations

$$\text{Volume of sodium hydroxide required } (V) = V - \frac{V_2 \times N_2}{N_1} \quad 10$$

$$\text{Percent purity of ammonium nitrate} = \frac{V - N_1 \times 8.005}{M} \quad 11$$

$$\text{Purity (on dry basis), percent by mass} = \frac{100 \times \text{purity as obtained above}}{100 \times \text{moisture percent}}$$

12

Where;

N_1 = normality of sodium hydroxide,

N_2 = normality of hydrochloric acid, and

M = mass of the sample (g).

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Annex M (normative)

Determination of determination of metals and compounds of metals other than alkali metals, calcium and iron

M.1 Procedure

- M.1.1** Treat 5 g of the material in a porcelain crucible with 1 ml of concentrated sulphuric acid.
- M.1.2** Drive off the excess acid by gentle heating and heat the residue strongly.
- M.1.3** Extract the residue with dilute hydrochloric acid and test qualitatively for the presence of metallic radicals (other than alkali metals, iron and calcium) like lead, copper zinc and manganese.
- M.1.4** If any metallic impurity is detected, estimate quantitatively by a standard colorimetric method

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Annex N (normative)

Determination of pyridine

N.1 Pyridine

Triturate a little powdered borax rapidly with about 5 g of the material in a mortar. If the odour of pyridine is detected, estimate the amount of pyridine present as follows

N.2 Procedure

N.2.1 Dissolve 100 g of the material in 300 ml of water in a 500-ml flask and add a few drops of methyl orange solution.

N.2.2 Fit the flask with a rubber stopper through which passes a tap-funnel and a glass outlet tube.

N.2.3 The outlet tube incorporating a bulb or other device to guard against a suck back, leads into a 500-ml conical flask containing sodium hydroxide solution and 4 ml of bromine, and dips below the surface of this solution.

N.2.4 Close the conical flask with a rubber stopper through which also passes a distilling trap attached to a water condenser.

N.2.5 Attach to the exit tube of the condenser a 250-ml conical flask as receiver containing a little water and fitted with guard tube of moistened glass beads.

N.2.6 Add 5 ml of 1 N sodium hydroxide solution to the ammonium nitrate solution through the tap funnel and distil the solution slowly into the hypobromite solution. When about 50 ml of the ammonium nitrate solution has been distilled, steam distil the hypobromite solution by warming the flask of hypobromite and continuing the distillation of the ammonium nitrate solution. Titrate the pyridine so collected in the receiver with 0.1 N hydrochloric acid using methyl orange as indicator

N.2.7 Calculate the factor to be used in calculating the amount of pyridine present from the results of titration of aqueous solutions, containing known quantities of pyridine.

Annex O (normative)

Determination of thiocyanate

O.1 Dissolve 5 g of the material in about 80 ml of water. Filter the solution, if necessary, and make it up to 100 ml in a Nessler cylinder. Prepare similarly a series of standard solutions containing 5 g of pure thiocyanate free ammonium nitrate and different known amounts of ammonium thiocyanate.

O.2 Add 2 ml of 10 percent ferric alum solution (containing 50 ml of concentrated nitric acid/litre) to the contents of each cylinder, with stirring.

O.3 Determine the amount of thiocyanate present in the material by comparing the colour of the solution under test with the standard solutions

Annex P (normative)

Determination of prussian blue and allied products

P.1 Warm 10 g, approximately, of the material with about 20 ml of 1 N nitric acid solution. If the resulting solution is coloured blue or contains blue coloured particles, estimate the prussianblue as follows.

P.2 Dissolve 100 g of the material in about 100 ml of water and acidify the solution with 20 ml of 1 N nitric acid. Heat for 5 minutes and then filter. Treat the residue of prussian blue on the filter paper with 100 ml of 1 N potassium hydroxide solution, added in consecutive small portions and make up the resulting filtrate, containing potassium ferrocyanide, to a volume of 200 ml.

P.3 Transfer a suitable aliquot portion of the ferrocyanide solution to a Nessler cylinder and add a 2 ml excess of 1 N nitric acid. Dilute the solution to 100 ml and add 5 ml of 10 percent ferric alum solution with stirring

P.4 Prepare similarly a series of standard solutions in Nessler cylinders, containing the same quantities of alkali, acid and ferric alum as in the first Nessler cylinder but with different known quantities of potassium ferrocyanide.

P.5 Determine the amount of prussian blue present in the material by comparing the colour of the solution under test with the standard solutions.

Annex Q (normative)

Determination of oil absorption

Q.1 Procedure

Q.1.1 10 g of sample to be tested is placed in a dry 50 ml beaker.

Q.1.2 Ten ml of diesel oil is added to the sample in the beaker and the contents are left for exactly 10 minutes.

Q.1.3 The contents of the beaker are then transferred to a porcelain buchner funnel (6 cm internal diameter) equipped with a Whatman No. 1 filter paper.

Q.1.4 A suction of 150 mm mercury is applied for three minutes.

Q.1.5 The oiled sample is then transferred to a previously weighed watch glass.

Q.1.6 The mass of the oiled sample is calculated from the combined masses of the sample and the watch glass. The percent oil absorption is calculated as follows.

Q.2 Calculation

$$\text{Oil absorption, percent by mass} = \frac{M_2 - M_1}{M_1} * 100$$

13

Where;

M₂ = mass in g of oiled sample, and

M₁ = mass in g of original sample

Annex R (normative)

Determination of bulk density

R.1 Apparatus

Assemble the apparatus as shown in Figure 1 (All dimensions in millimetres)

The measuring cylinder A shall be of 20 ml capacity, base of the measuring cylinder shall be ground flat and the distance between the flat-ground part of the base of the measuring cylinder A and the rubber base pad B, when the measuring cylinder A is raised to the full height, shall be $25 \text{ mm} \pm 2 \text{ mm}$.

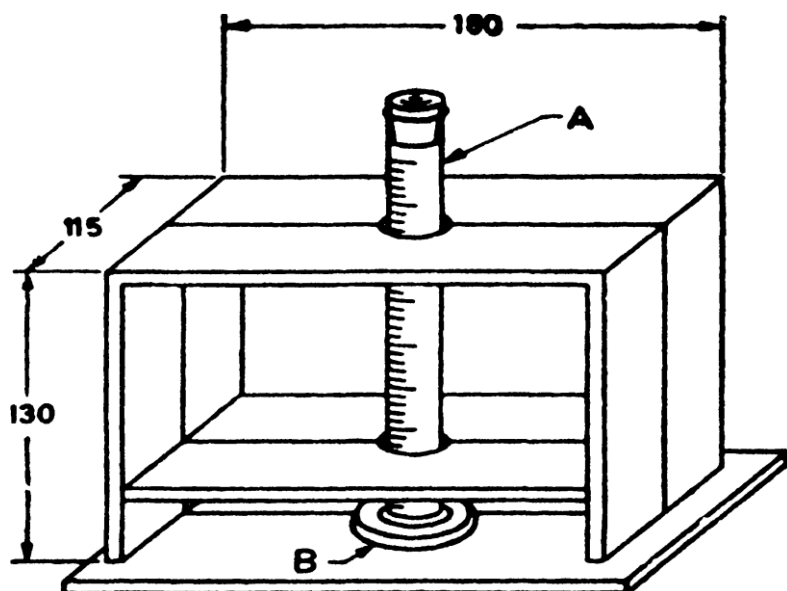


Figure 1 — Apparatus for determination of bulk density

The rubber base pad B shall have a shore hardness of 42 to 50.

Pans of the balance shall be at least 10 cm in diameter and the balance shall be sensitive to less than 0.1 g.

R.2 Procedure

Place 50 g of the material in the graduated cylinder fitted in the apparatus and tap it until a constant volume of the material is obtained. Note down the volume of the material

R.3 Calculation

$$\text{Bulk density, g/cm}^3 = \frac{M}{V}$$

14

Where;

M = mass in g of the material taken for the test, and

V = volume of the material in the cylinder.

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Annex S (normative)

Determination of total nitrogen

S.1 Outline of the method

Ammonia is distilled off from an alkaline solution and absorbed in standard hydrochloric acid. The excess of acid is titrated with standard hydroxide solution.

S.2 Apparatus

Ammonia distillation apparatus as shown in Figure 2

S.3 Reagents

- 1) **Sodium Hydroxide solution** — 30 percent (m/v).
- 2) **Standard hydrochloric acid** — 0.5 N.
- 3) **Standard Sodium Hydroxide solution** — 0.5 N.
- 4) **Mixed Methyl Red Indicator solution** — Dissolve 0.1 g of methyl red in 50 ml of ethyl alcohol and 0.05 g of methylene blue in 50 ml of ethyl alcohol and mix the two solutions
- 5) **Devarda's Alloy** — Containing 45 parts of aluminium, 50 parts of copper and 5 parts of zinc

S.4 Procedure

S.4.1 Weigh accurately 2.0 g of the prepared sample and transfer to the Kjeldahl flask, adding about 150 ml of water and a few porcelain beads. Add 6 to 7 g Devarda's alloy. Connect the apparatus as shown in Figure. 2.

S.4.2 Measure into the beaker exactly 125 ml of standard hydrochloric acid and add 4 to 5 drops of the mixed indicator solution.

S.4.3 Place the beaker so that the end of the delivery tube is below the surface of the acid. Pour about 25 ml of sodium hydroxide solution (30 percent) into the tap funnel. Slowly drop the sodium hydroxide solution into the Kjeldahl flask by opening the tap of the funnel. Close the tap after delivering the entire quantity of sodium hydroxide.

S.4.4 Start heating the Kjeldahl flask and circulate water in the condenser. After distilling about 150 ml into the receiver, remove the beaker and titrate the excess acid using 0.5 N sodium hydroxide solution. The end point is orange to green. Suggested indicator solution for the above determination is methyl red and bromocresol green solution. It is prepared by dissolving 0.1 g of methyl red in 50 ml of ethyl alcohol and mixing the two solutions.

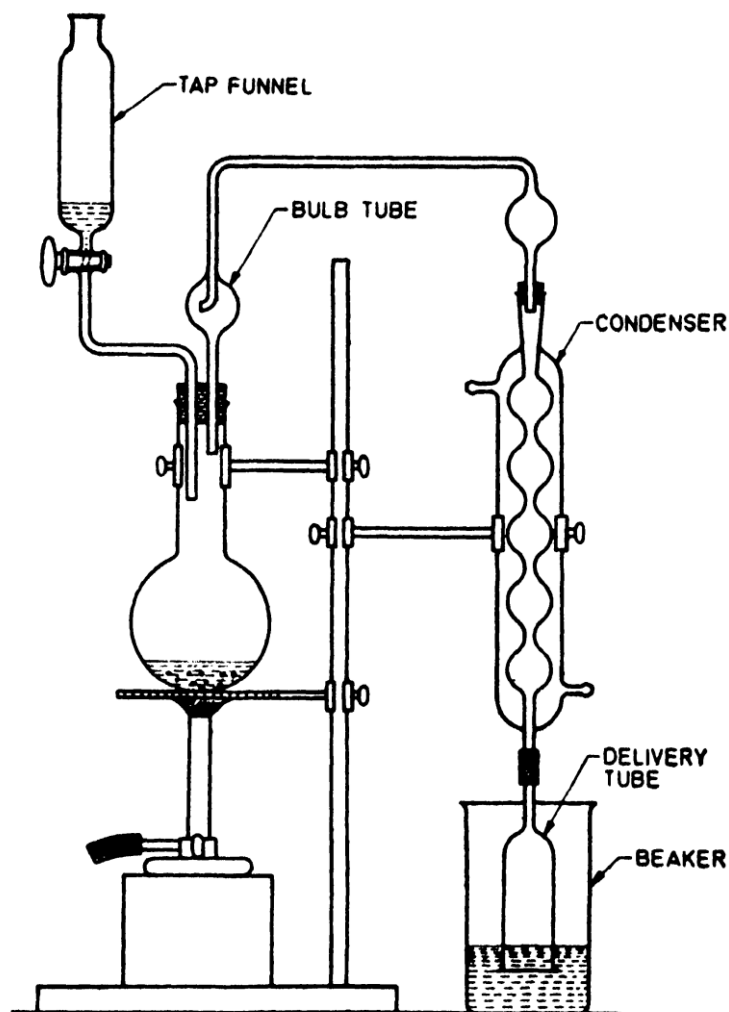


Figure 2 — Apparatus for determination of nitrogen

S.4 Calculation

$$\text{Total nitrogen (as N), percent by mass} = \frac{1.401 * (V_1N_1 - V_2N_2)}{M}$$

15

Where;

V_1 = volume in ml of standard hydrochloric acid used,

N_1 = normality of standard hydrochloric acid,

V_2 = volume in ml of standard sodium hydroxide solution,

N_2 = normality of standard sodium hydroxide solution, and

M = mass in g of the material taken for the test.

Annex T (normative)

Determination of sulphates

T.1 Outline of the Method

Sulphates are determined by comparing the turbidity produced with barium sulphate solution, against a standard turbidity in Nessler cylinders

T.2 Apparatus

- 1) **Nessler Cylinders** — 50-ml capacity

T.3 Reagents

- 1) Sodium Carbonate — Solid
- 2) Dilute Hydrochloric Acid — Approximately 2 N
- 3) Standard Sulphuric Acid — 0.002 N
- 4) Barium Sulphate Reagent — Dissolve 2.0 g of barium chloride in 75 ml of water, add 20 ml of 95 percent (v/v) ethyl alcohol and 5 ml of standard sulphuric acid. It shall be prepared afresh

T.4 Procedure

T.4.1 Weigh accurately 10 g of the prepared sample into a porcelain dish and dissolve in 40 ml of water. Add 0.05 g of sodium carbonate and heat on a sand-bath at about 120 °C to volatilize water and then heat over a very low flame until the evolution of white fumes has ceased. Finally heat the dish at 600°C in a furnace for 15 min. Cool, dissolve the residue in 10 ml of water, add 1 ml. of dilute hydrochloric acid and filter through a Whatman No. 41 filter paper.

T.4.2 Transfer the filtrate to a Nessler cylinder, add 10 ml of water and 1 ml of barium sulphate reagent. Mix well and allow to stand for 5 min.

T.4.3 Prepare a series of sulphate standard turbidities with different known volumes of standard sulphuric acid and 1 ml of barium sulphate reagent in the similar manner and compare the turbidities.

T.4.4 Calculate the sulphate content as ammonium sulphate in the sulphate standard in which the turbidity most closely matches with that in the solution of the material.

Bibliography

- [1] IS 4668 – 1985 Specification for Ammonium Nitrate for explosives

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